

nucleoside residue in the axial position on it. The principal difference, apart from the base, lies in the disposition of the tetrahydropyran ring with respect to the O(2')-C(2'') bond. This amounts to an interchange of the C(3'') and O(1'') atoms and a rotation of 90° about C(2'') → O(2'). We may assume that the tetrahydropyran ring in the more laevorotatory isomer of ThPU will have a similar disposition to that in ThPA(-) so that ThPU(-) would have the molecular structure shown in Fig. 6. A comparison of the pictures of the two isomers (Fig. 6) shows that ThPU(+) can form an internal hydrogen bond [O(3')-H(O3')-O(1'')] with very little distortion of the molecule, the distance O(3')-O(1'') in the crystal being only 3.4 Å. By contrast, in ThPU(-), a large rotation would be necessary to bring O(3') and O(1'') within hydrogen-bonding distance, and such rotation would lead to unacceptably short atomic contact distances, especially between H(12'') and H(11') or O(3'). Thus in a non-polar solvent, ThPU(+) is able to present a less polar exterior than ThPU(-) with the resultant observed differences in their physical properties.

We wish to thank Mr K. V. Deugau for crystallizing the compound and the National Research Council of Canada for financial support.

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## The Crystal Structure of Hexalead Pentaborate, 6PbO · 5B<sub>2</sub>O<sub>3</sub>

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(Received 2 April 1973; accepted 30 May 1973)

Hexalead pentaborate, 6PbO · 5B<sub>2</sub>O<sub>3</sub>, is triclinic with the following unit-cell dimensions at 22°C:  $a = 6.833 \pm 0.002$ ,  $b = 7.010 \pm 0.002$ ,  $c = 11.125 \pm 0.002$  Å,  $\alpha = 85.89 \pm 0.03$ ,  $\beta = 102.00 \pm 0.02$ ,  $\gamma = 119.56 \pm 0.02^\circ$ . The space group is  $P\bar{1}$  with one formula unit in the cell and a calculated density of 6.183 g cm<sup>-3</sup>. Three-dimensional X-ray intensity data were measured with a Picker automatic single-crystal diffractometer and Mo K $\alpha$  radiation. The structure was determined by the heavy-atom method. A full-matrix least-squares refinement resulted in an  $R$  value of 0.065. The structure contains isolated B<sub>10</sub>O<sub>21</sub><sup>12-</sup> polyions. These ions are built up from two diborate groups linked by two BO<sub>3</sub> triangles. The lead atoms are covalently bonded to oxygen, as evidenced by the short lead-oxygen distances (2.23 Å and upwards) as well as the asymmetric coordination of the lead atoms.

#### Introduction

The system lead oxide-boron oxide exhibits two phases of unusual stoichiometry. Geller & Bunting (1937) reported the existence of 1:2, 5:4, 2:1 and 4:1 phases in this system. Bauer & Plötscher (1967) claimed to

have found in addition 1:1 and 9:8 phases. Crystalline compounds with the 5:4 and 9:8 stoichiometric ratios of metal oxide to boron oxide have not been found for other borate systems. The present investigation was undertaken to study the nature of the phase which Bauer & Plötscher believed to be a 9PbO · 8B<sub>2</sub>O<sub>3</sub> com-

Table 1. Final observed and calculated structure factors. The columns are I, 10F<sub>o</sub> and 10F<sub>c</sub> respectively.

Table with multiple columns containing numerical data for structure factors. The columns are labeled I, 10F<sub>o</sub>, and 10F<sub>c</sub>. The data is organized in a grid-like format with various numerical values and some text labels like 'h-k-l' and 'I' scattered throughout the rows.



Table 2 (cont.)

	x	y	z	U
O(1)	-1679 (31)	3211 (29)	2012 (16)	112 (32)
O(2)	-2780 (31)	-447 (29)	2354 (16)	125 (32)
O(3)	1187 (25)	2324 (23)	2944 (13)	17 (24)
O(4)	5121 (28)	4806 (25)	2888 (14)	64 (27)
O(5)	3257 (30)	5683 (27)	4196 (15)	104 (31)
O(6)	2254 (32)	7804 (30)	5329 (17)	146 (34)
O(7)	321 (27)	6627 (24)	3234 (14)	47 (26)
O(8)	2273 (28)	5786 (26)	1981 (15)	78 (28)
O(9)	-590 (31)	6795 (28)	1080 (16)	112 (31)
O(10)	3042 (29)	9347 (27)	540 (15)	93 (30)
O(11)	0	0	0	361 (79)
B(1)	2992 (45)	4694 (41)	2976 (23)	88 (43)
B(2)	90 (44)	5657 (41)	2026 (23)	80 (42)
B(3)	-1059 (51)	1629 (47)	2417 (26)	141 (50)
B(4)	1873 (39)	6661 (36)	4262 (20)	31 (36)
B(5)	940 (37)	8737 (34)	563 (19)	12 (34)

factors, were those of *International Tables for X-ray Crystallography* (1962). The atomic scattering factors for neutral Pb were obtained from Hanson, Herman, Lea & Skillman (1964) and were corrected for dispersion. The imaginary dispersion terms (Cromer, 1965) were also used in the refinement. The vibrations of boron and oxygen were assumed to be isotropic. A weighting scheme based on the statistical counting errors, compounded with errors assumed to be 1% of the observed intensity, was used. The structure was refined to an *R* value of 0.065 (0.081 for the weighted *R* value), assuming the presence of a centre of symmetry, *i.e.* space group  $P\bar{1}$ . The ratio of maximum

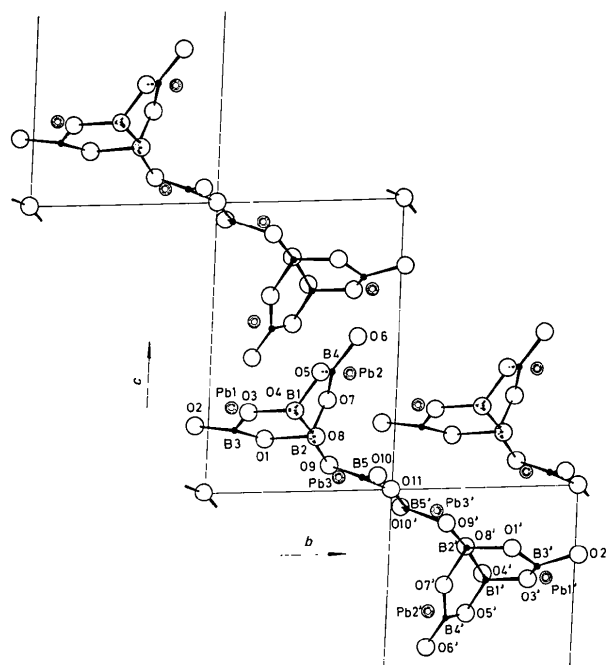


Fig. 1. View of the structure of  $6\text{PbO} \cdot 5\text{B}_2\text{O}_3$  along the *a* axis. The concentric circles represent lead, the small filled circles boron and the remaining circles oxygen. The atom numbering corresponds to the numbering used in the text.

shift to standard deviation during the last cycle was 0.08. (An attempt was made to refine the structure without the restriction of a centre of symmetry, but the refinement did not converge properly.) Observed and calculated structure factors are given in Table 1. The final atomic coordinates and thermal parameters are given in Table 2.

### Discussion of the structure

The anion structure of hexalead pentaborate consists of  $\text{B}_{10}\text{O}_{21}^{12-}$  isolated units (Figs. 1 and 2). This new anion type is the largest isolated borate anion found so far in an anhydrous borate. The anion may be described as two diborate groups connected to each other by a chain section consisting of two  $\text{BO}_3$ -triangles. The diborate group was first found by Morimoto (1956) in the mineral borax and has since been found in several borates, both anhydrous and hydrated. Each diborate subunit in the hexalead pentaborate anion is terminated by three oxygen atoms bonded only to one boron, not counting bonds to the lead atoms. Isolated diborate groups have previously only been found with hydroxyl units as terminal atoms. Boron-oxygen bond lengths and bond angles are given Table 3. The values are normal within the error limits. As usual the longest boron-oxygen distances are found for the tetrahedrally coordinated boron atoms.

The lead coordination has noteworthy features. The three crystallographically different lead atoms are all coordinated by three or four oxygens at a short distance (2.23–2.55 Å) and oxygens at longer distances, upwards of 2.76 Å. All the short lead-oxygen distances of a given lead atom fall within the same hemisphere around lead. Moreover the short lead-oxygen distances, with two exceptions [O(8) and O(7)], involve terminal oxygen atoms, *i.e.* oxygens bonded to only one boron. One exception is the lead-oxygen bond to the bridging O(8) atom. This atom, however, is engaged in the two relatively weak boron-oxygen bonds of a pair of  $\text{BO}_4$ -tetrahedra.

Table 3. Interatomic distances and bond angles

B(1)-O(3)	1.50 (3) Å	Pb(1)-O(6)	2.28 (2) Å
B(1)-O(4)	1.44 (4)	Pb(1)-O(4)	2.29 (2)
B(1)-O(5)	1.50 (3)	Pb(1)-O(1)	2.55 (2)
B(1)-O(8)	1.44 (3)	Pb(1)-O(2)	2.76 (3)
B(2)-O(1)	1.53 (3)	Pb(1)-O(3)	2.85 (2)
B(2)-O(7)	1.49 (3)	Pb(1)-O(10)	2.94 (2)
B(2)-O(8)	1.46 (4)	Pb(2)-O(4)	2.23 (1)
B(2)-O(9)	1.41 (3)	Pb(2)-O(2)	2.39 (2)
B(3)-O(1)	1.39 (4)	Pb(2)-O(7)	2.44 (2)
B(3)-O(2)	1.34 (3)	Pb(2)-O(5)	2.79 (2)
B(3)-O(3)	1.36 (3)	Pb(2)-O(5)	2.81 (2)
B(4)-O(5)	1.43 (4)	Pb(2)-O(6)	2.93 (2)
B(4)-O(6)	1.39 (3)	Pb(3)-O(10)	2.30 (2)
B(4)-O(7)	1.38 (3)	Pb(3)-O(2)	2.33 (2)
B(5)-O(9)	1.42 (3)	Pb(3)-O(8)	2.38 (2)
B(5)-O(10)	1.29 (3)	Pb(3)-O(10)	2.48 (2)
B(5)-O(11)	1.38 (3)	Pb(3)-O(4)	3.00 (2)
		Pb(3)-O(1)	3.19 (2)
		Pb(3)-O(9)	3.26 (1)

Table 3 (cont.)

O(3)—B(1)—O(4)	108 (2) <sup>o</sup>	O(2)—Pb(3)—O(10')	83 (1) <sup>o</sup>
O(3)—B(1)—O(5)	107 (2)	O(8)—Pb(3)—O(10')	144 (1)
O(3)—B(1)—O(8)	109 (2)	O(10')—Pb(3)—O(4)	146 (< 1)
O(4)—B(1)—O(5)	109 (2)		
O(4)—B(1)—O(8)	112 (2)	B(2)—O(1)—B(3)	122 (2) <sup>o</sup>
O(5)—B(1)—O(8)	111 (3)	B(1)—O(3)—B(3)	122 (2)
O(7)—B(2)—O(1)	107 (2)	B(1)—O(5)—B(4)	118 (2)
O(7)—B(2)—O(8)	108 (2)	B(4)—O(7)—B(2)	122 (2)
O(7)—B(2)—O(9)	108 (2)	B(1)—O(8)—B(2)	113 (2)
O(8)—B(2)—O(1)	107 (2)	B(2)—O(9)—B(5)	125 (2)
O(8)—B(2)—O(9)	113 (2)		
O(1)—B(2)—O(9)	113 (2)	O(6)—Pb(1)—O(4)	97 (1)
O(1)—B(3)—O(2)	116 (3)	O(6)—Pb(1)—O(1)	78 (1)
O(1)—B(3)—O(3)	118 (2)	O(4)—Pb(1)—O(1)	84 (1)
O(2)—B(3)—O(3)	126 (3)	O(4)—Pb(2)—O(2)	80 (1)
O(5)—B(4)—O(6)	119 (2)	O(4)—Pb(2)—O(7)	82 (< 1)
O(5)—B(4)—O(7)	120 (2)	O(2)—Pb(2)—O(7)	96 (1)
O(6)—B(4)—O(7)	121 (3)	O(10)—Pb(3)—O(2)	83 (1)
O(9)—B(5)—O(10)	126 (2)	O(10)—Pb(3)—O(8)	72 (1)
O(9)—B(5)—O(11)	115 (2)	O(10)—Pb(3)—O(10')	73 (1)
O(10)—B(5)—O(11)	119 (2)	O(2)—Pb(3)—O(8)	81 (1)

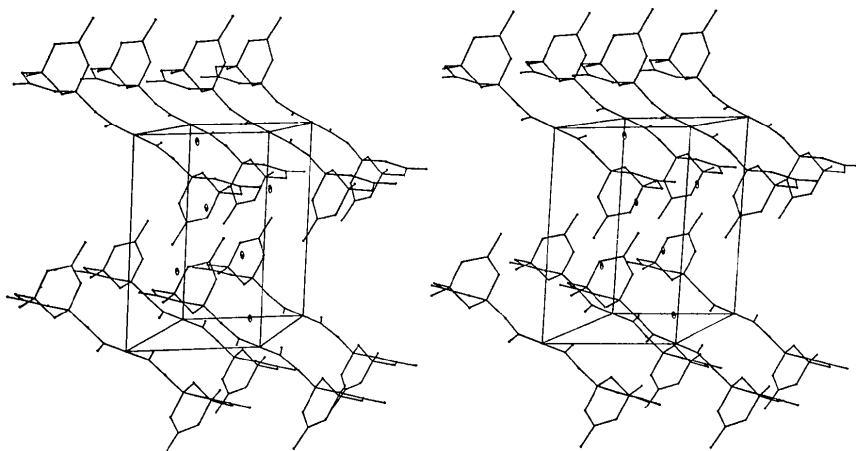


Fig. 2. A stereo pair showing the packing of the structure. The boron-oxygen bonding and the lead atoms are shown.

These characteristics point to the covalent nature of the lead-oxygen bonding. The shortest lead-oxygen distance of 2.23 Å has a value similar to the short lead-oxygen distance of 2.18 Å in lead oxide (Leciejewicz, 1961; Kay, 1961). An ionic lead-oxygen distance, in comparison, would be about 2.60 Å (Pauling, 1960). The asymmetric coordination around the lead atoms in lead oxide, caused by the two 6s electrons, is also retained in the borate compound studied here, emphasizing the similarity with lead oxide in the nature of the bonding.

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